



basic education

Department:
Basic Education
REPUBLIC OF SOUTH AFRICA

**NATIONAL
SENIOR CERTIFICATE**

GRADE 12

PHYSICAL SCIENCES: CHEMISTRY (P2)

FEBRUARY/MARCH 2011

MARKS: 150

TIME: 3 hours

This question paper consists of 16 pages and 4 data sheets.

INSTRUCTIONS AND INFORMATION

1. Write your centre number and examination number in the appropriate spaces on the ANSWER BOOK.
2. Answer ALL the questions in the ANSWER BOOK.
3. This question paper consists of TWO sections:

SECTION A (25)
SECTION B (125)
4. You may use a non-programmable calculator.
5. You may use appropriate mathematical instruments.
6. Number the answers correctly according to the numbering system used in this question paper.
7. Data sheets and a periodic table are attached for your use.
8. Give brief motivations, discussions, et cetera where required.

SECTION A**QUESTION 1: ONE-WORD ITEMS**

Give ONE word/term for each of the following descriptions. Write only the word/term next to the question number (1.1 – 1.5) in the ANSWER BOOK.

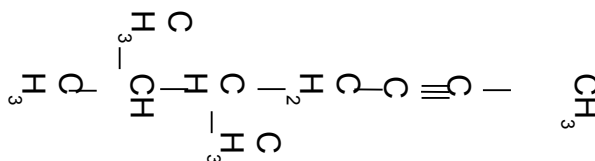
- 1.1 The resistance that a fluid offers to flow (1)
- 1.2 A chemical substance that provides an alternative path of lower activation energy for a chemical reaction (1)
- 1.3 The acid produced during the contact process (1)
- 1.4 The name of the electrode in a galvanic (voltaic) cell at which oxidation takes place (1)
- 1.5 The industrial process for the production of ammonia (1)
- [5]**

QUESTION 2: MULTIPLE-CHOICE QUESTIONS

Four options are provided as possible answers to the following questions. Each question has only ONE correct answer. Write only the letter (A – D) next to the question number (2.1 – 2.10) in the ANSWER BOOK.

- 2.1 Which ONE of the following pairs of compounds correctly represents the products formed during the COMPLETE combustion of octane?
- A CO and H₂O
- B CO and H₂
- C CO₂ and H₂
- D CO₂ and H₂O (2)
- 2.2 Which ONE of the following pairs of reactants can be used to prepare the ester ethyl methanoate in the laboratory?
- A Ethane and methanoic acid
- B Methanol and ethanoic acid
- C Ethanol and methanoic acid
- D Ethene and methanol (2)

2.3 The structural formula of an organic compound is given below.



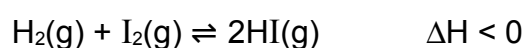
The IUPAC name of this compound is ...

- A 2,3-dimethylhept-5-yne.
- B 5,6-dimethylhept-2-yne.
- C 2,3-methylhept-2-yne.
- D 5,6-dimethylhept-3-yne. (2)

2.4 The type of compound formed when but-1-ene reacts with water in the presence of a suitable catalyst is a/an ...

- A alcohol.
- B alkane.
- C haloalkane.
- D ester. (2)

2.5 The equation below represents a chemical reaction at equilibrium in a closed container.



Which ONE of the following changes will increase the yield of HI(g) in the above reaction?

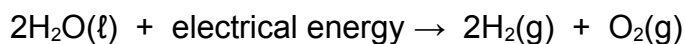
- A Increase the temperature
- B Decrease the temperature
- C Increase the pressure by decreasing the volume
- D Decrease the pressure by increasing the volume (2)

2.6 A chemical reaction reaches equilibrium. Which ONE of the following statements regarding this equilibrium is TRUE?

- A The concentrations of the individual reactants and products are constant.
- B The concentrations of the individual reactants and products are equal.
- C The concentrations of the individual reactants are zero.
- D The concentrations of the individual products increase until the reaction stops.

(2)

2.7 The net (overall) cell reaction taking place in a certain cell is represented as follows:



Which ONE of the following statements best describes this cell?

The cell is a/an ...

- A electrolytic cell in which an exothermic reaction occurs.
- B electrolytic cell in which an endothermic reaction occurs.
- C galvanic (voltaic) cell in which an exothermic reaction occurs.
- D galvanic (voltaic) cell in which an endothermic reaction occurs.

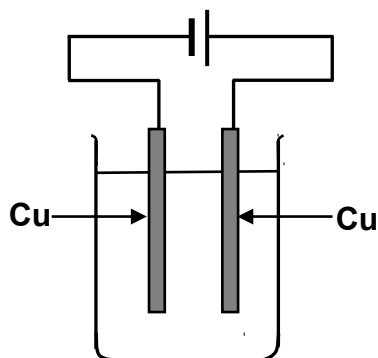
(2)

2.8 When the net (overall) cell reaction in a galvanic (voltaic) cell reaches equilibrium, the emf of the cell is equal to ...

- A +2,00 V.
- B +1,00 V.
- C 0,00 V.
- D -1,00 V.

(2)

- 2.9 Copper is purified through electrolysis as represented in the simplified diagram below.



Which ONE of the following statements is CORRECT for this process?

- A Cu is oxidised at the negative electrode.
- B Cu is reduced at the positive electrode.
- C Cu^{2+} ions are reduced at the positive electrode.
- D Cu^{2+} ions are reduced at the negative electrode. (2)

- 2.10 The major products formed in the chlor-alkali industry are ...

- A chlorine gas and sodium hydroxide.
 - B chlorine gas and sodium chloride.
 - C hydrogen chloride gas and sodium hydroxide.
 - D chlorine gas and hydrogen chloride gas. (2)
- [20]

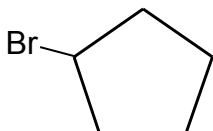
TOTAL SECTION A: 25

SECTION B**INSTRUCTIONS**

1. Start EACH question on a NEW page.
2. Leave ONE line between two subquestions, for example between QUESTION 3.1 and QUESTION 3.2.
3. Show the formulae and substitutions in ALL calculations.
4. Round off your answers to TWO decimal places, where applicable.

QUESTION 3 (Start on a new page.)

Millions of organic compounds are known to date. Four of these compounds, represented by the letters **P**, **Q**, **R** and **S**, are shown in the table below.

P	methanal	Q	<pre> H H O H H - C - C - C - C - H H H H H </pre>
R		S	<pre> H H H H H - C - C - C - C - H H O H H H </pre>

- 3.1 Write down the following:
- 3.1.1 Structural formula of the functional group of **P** (2)
 - 3.1.2 Homologous series to which **Q** belongs (1)
 - 3.1.3 Structural formula of an isomer of **Q** (2)
 - 3.1.4 IUPAC name of **R** (2)
- 3.2 **S** represents an alcohol. Classify this alcohol as primary, secondary or tertiary. (1)

[8]

QUESTION 4 (Start on a new page.)

Knowledge of boiling points can be used to identify chemical compounds. The boiling points of four organic compounds, represented by the letters **A**, **B**, **C** and **D**, are given in the table below.

	COMPOUND	BOILING POINT (°C)
A	Propane	-42
B	Pentane	36
C	2-methylbutane	27,8
D	Pentan-1-ol	137

- 4.1 Define the term *boiling point*. (2)
- 4.2 Which ONE of **A** or **B** has the higher vapour pressure? (1)
- 4.3 An unknown STRAIGHT CHAIN ALKANE has a boiling point of $-0,5\text{ }^{\circ}\text{C}$. Use the information in the table to identify this alkane and write down its IUPAC name. (2)
- 4.4 **B** and **C** are structural isomers.
- 4.4.1 Define the term *structural isomer*. (2)
- 4.4.2 Explain why **B** has a higher boiling point than **C**. Refer to structure, intermolecular forces and energy in your explanation. (3)
- 4.5 Explain the difference in the boiling points of **B** and **D**. Refer to intermolecular forces and energy in your explanation. (4)
- [14]**

QUESTION 5 (Start on a new page.)

- 5.1 Prop-1-ene, an UNSATURATED hydrocarbon, and compound **X**, a SATURATED hydrocarbon, react with chlorine, as represented by the incomplete equations below.



- 5.1.1 Give a reason why prop-1-ene is classified as unsaturated. (1)
- 5.1.2 What type of reaction (ADDITION or SUBSTITUTION) takes place in the following:
- (a) Reaction I (1)
- (b) Reaction II (1)
- 5.1.3 Write down the structural formula of the product formed in Reaction I. (2)
- 5.1.4 Write down the reaction condition necessary for Reaction II to take place. (1)
- 5.1.5 Write down the IUPAC name of reactant **X**. (1)
- 5.1.6 Write down the name or formula of product **Y**. (1)
- 5.2 2-chlorobutane can either undergo ELIMINATION or SUBSTITUTION in the presence of a strong base such as sodium hydroxide.
- 5.2.1 Which reaction will preferably take place when 2-chlorobutane is heated in the presence of CONCENTRATED sodium hydroxide in ethanol? Write down only SUBSTITUTION or ELIMINATION. (1)
- 5.2.2 Write down the IUPAC name of the major organic compound formed in QUESTION 5.2.1. (2)
- 5.2.3 Use structural formulae to write down a balanced equation for the reaction that takes place when 2-chlorobutane reacts with a DILUTE sodium hydroxide solution. (6)
- 5.2.4 Write down the name of the type of substitution reaction that takes place in QUESTION 5.2.3. (1)
- 5.3 Haloalkanes are used in insecticides (insect killers).
- 5.3.1 Write down ONE POSITIVE impact of insecticides on human development. (2)
- 5.3.2 Write down ONE NEGATIVE impact of insecticides on humans. (2)

QUESTION 6 (Start on a new page.)

Learners perform three investigations (**A**, **B** and **C**) to study three factors which affect the rate of chemical reactions. They use the reaction between solid calcium carbonate (CaCO_3) and excess hydrochloric acid (HCl) solution, represented by the balanced equation below, in all three investigations.



EXCESS HYDROCHLORIC ACID is used and the calcium carbonate is COMPLETELY COVERED in all the investigations.

6.1 INVESTIGATION A:

The learners conduct two experiments using the conditions as shown in the table below.

	Mass of CaCO_3 (g)	State of CaCO_3	Concentration of HCl ($\text{mol}\cdot\text{dm}^{-3}$)	Temperature of HCl ($^{\circ}\text{C}$)
Experiment 1	2	powder	0,2	25
Experiment 2	2	lumps	0,2	25

6.1.1 Which factor influencing reaction rate is investigated? (1)

6.1.2 Write down an INVESTIGATIVE QUESTION for this investigation. (2)

6.1.3 The learners now repeat **Experiment 1**, but use 4 g of calcium carbonate in excess acid, instead of 2 g. They find that the rate of the reaction INCREASES.

Give a reason why the rate increases. (2)

6.2 INVESTIGATION B:

The learners conduct two experiments using the conditions as shown in the table below.

	Mass of CaCO ₃ (g)	State of CaCO ₃	Concentration of HCl (mol·dm ⁻³)	Temperature of HCl (°C)
Experiment 3	2	lumps	0,2	25
Experiment 4	2	lumps	1,0	25

- 6.2.1 Identify the independent variable in this investigation. (1)
- 6.2.2 Write down a hypothesis for this investigation. (2)
- 6.2.3 Is it fair to compare results obtained in **Experiment 3** with that in **Experiment 4**? Give a reason for the answer. (2)
- 6.2.4 The reactions in **Experiments 3** and **4** both run to completion. How will the yield of CO₂(g) in **Experiment 3** compare to that in **Experiment 4**? Write down only LARGER THAN, SMALLER THAN or EQUAL TO and give a reason for the answer. (2)

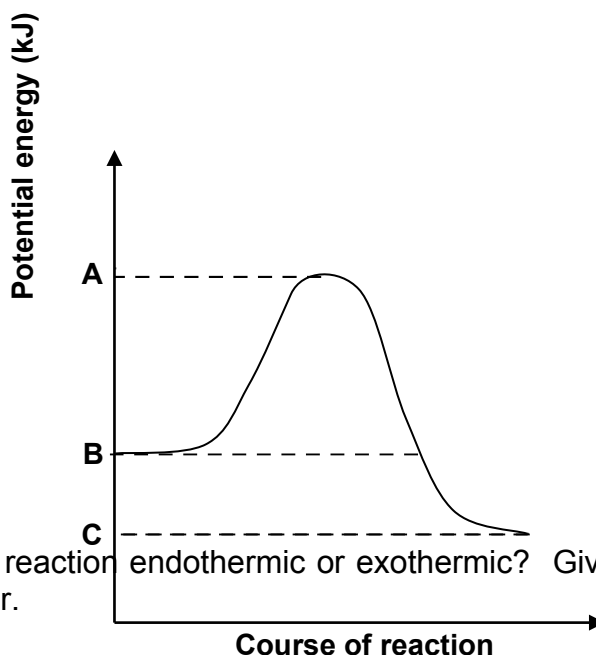
6.3 INVESTIGATION C:

The learners conduct two experiments using the conditions as shown in the table below.

	Mass of CaCO ₃ (g)	State of CaCO ₃	Concentration of HCl (mol·dm ⁻³)	Temperature of HCl (°C)
Experiment 5	4	powder	0,2	25
Experiment 6	4	powder	0,2	35

- 6.3.1 How does the average kinetic energy of the particles in the reaction in **Experiment 5** compare to that in **Experiment 6**? Write down only HIGHER THAN, LOWER THAN or EQUAL TO. (1)
- 6.3.2 On the same set of axes, draw sketch graphs of the number of molecules versus the kinetic energy (Maxwell-Boltzmann distribution curves) for each of **Experiment 5** and **Experiment 6**.
 • Label the axes.
 • Clearly label each graph as **Experiment 5** or **Experiment 6**. (3)

- 6.4 The graph below shows changes in the potential energy for the reaction between calcium carbonate and hydrochloric acid.



- 6.4.1 Is this reaction endothermic or exothermic? Give a reason for the answer. (2)
- 6.4.2 Use the relevant energy values, **A**, **B** and **C**, to write down an expression for each of the following:
- (a) The energy of the activated complex (1)
- (b) ΔH for the forward reaction (1)
- [20]**

QUESTION 7 (Start on a new page.)

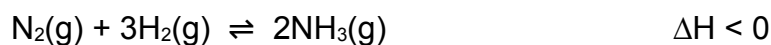
Fertilisers allow farmers to grow crops in the same soil year after year. However, environmental problems, such as eutrophication, are associated with the application of fertilisers.

- 7.1 State ONE PRECAUTION that a maize farmer can take to prevent eutrophication. (1)

Nitric acid is an important reactant in the production of ammonium nitrate, a nitrogen-based fertiliser.

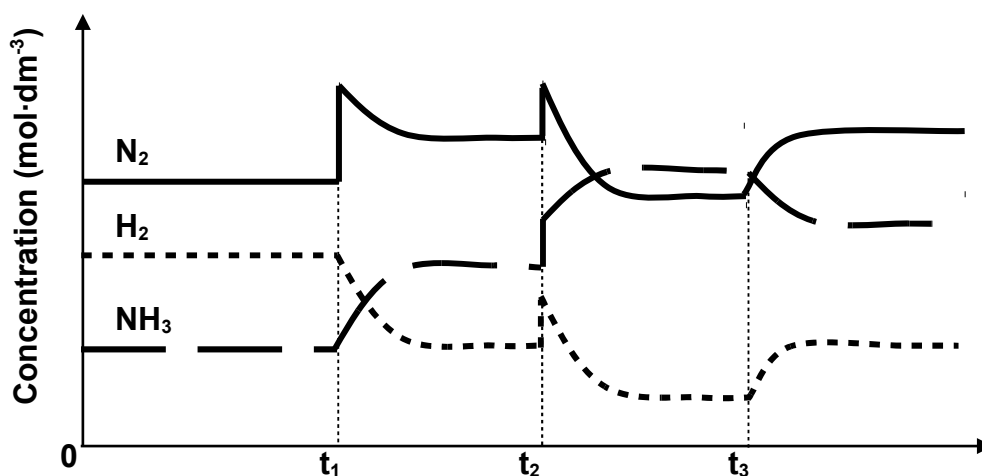
- 7.2 Write down the name of the industrial process for the production of nitric acid. (1)
- 7.3 Write down a balanced equation for the preparation of ammonium nitrate from nitric acid. (3)

A fertiliser company produces ammonia on a large scale at a temperature of 450 °C. The balanced equation below represents the reaction that takes place in a sealed container.



To meet an increased demand for fertiliser, the management of the company instructs their engineer to make the necessary adjustments to increase the yield of ammonia.

In a trial run on a small scale in the laboratory, the engineer makes adjustments to the TEMPERATURE, PRESSURE and CONCENTRATION of the equilibrium mixture. The graphs below represent the results obtained.



7.4 Identify the changes made to the equilibrium mixture at each of the following times:

7.4.1 t_1 (2)

7.4.2 t_2 (2)

7.4.3 t_3 (2)

Time (minutes)

At which of the above time(s) did the change made to the reaction mixture lead to a higher yield of ammonia? Write down only t_1 and/or t_2 and/or t_3 .

7.5 (2)

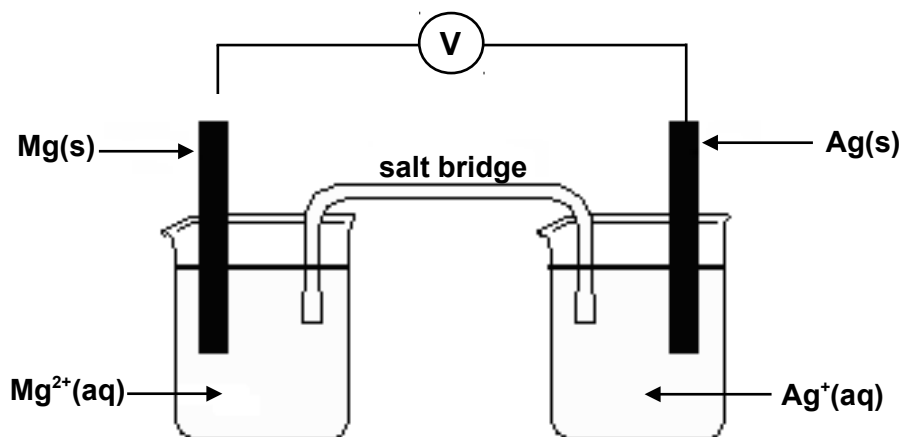
7.6 The engineer now injects 5 mol N_2 and 5 mol H_2 into a 5 dm³ sealed empty container. Equilibrium is reached at 450 °C. Upon analysis of the equilibrium mixture, he finds that the mass of NH_3 is 20,4 g.

Calculate the value of the equilibrium constant (K_c) at 450 °C (9)

[22]

QUESTION 8 (Start on a new page.)

The diagram below represents a galvanic (voltaic) cell functioning under standard conditions with magnesium and silver as electrodes. A voltmeter connected across the electrodes shows an initial reading of 3,17 V.

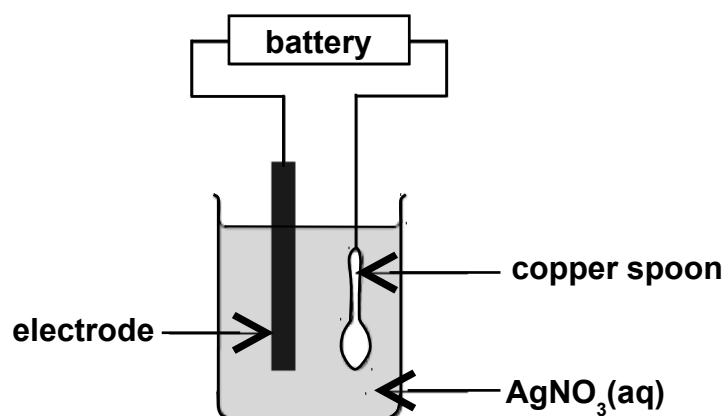


- 8.1 State the energy conversion that takes place in this cell. (2)
- 8.2 State TWO standard conditions under which this cell operates. (2)
- 8.3 Identify the anode of this cell. Refer to the relative strength of reducing agents to explain how you arrived at the answer. (3)
- 8.4 Write down the cell notation (symbolic notation) of this cell. (3)
- 8.5 Write down the balanced equation for the net (overall) cell reaction that takes place in this cell. Omit the spectator ions. (3)
- 8.6 How will an increase in the concentration of the Ag^+ ions influence the current that the cell delivers? Write down only INCREASES, DECREASES or REMAINS THE SAME and explain the answer. (3)

(3)
[16]

QUESTION 9 (Start on a new page.)

Electroplating is one of the uses of electrolysis. The diagram below shows an electrolytic cell that can be used to plate a copper spoon with silver.



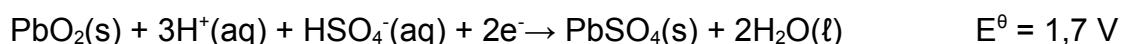
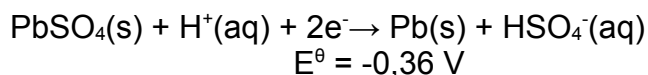
- 9.1 Define the term *oxidation* in terms of electron transfer. (2)
- 9.2 What type of half-reaction takes place at the copper spoon? Write down only OXIDATION or REDUCTION. (1)
- 9.3 Write down a half-reaction that explains the change that occurs on the surface of the copper spoon during electrolysis. (2)
- 9.4 Name the metal that is labelled 'electrode'. (1)
- 9.5 Give a reason why the concentration of the $\text{AgNO}_3(\text{aq})$ remains constant during electrolysis. (2)

[8]

QUESTION 10 (Start on a new page.)

A lead-acid battery (car battery) consists of six cells and has a battery capacity of 20 A·h.

The half-reactions that take place in each cell and their respective standard reduction potentials are represented below:



- 10.1 Are car batteries primary or secondary batteries? (1)
- 10.2 Write down the equation for the net (overall) cell reaction that takes place in each cell of this battery. (3)
- 10.3 Calculate the emf of the BATTERY, consisting of six cells, under standard conditions. (5)
- 10.4 Calculate the maximum time that this battery will be able to supply a constant current of 5 A to an appliance connected to it. Assume that the capacity of the battery remains constant. (4)
- 10.5 State TWO environmental risks associated with the irresponsible disposal of lead-acid batteries. (2)

[15]

TOTAL SECTION B: 125
GRAND TOTAL: 150

**DATA FOR PHYSICAL SCIENCES GRADE 12
PAPER 2 (CHEMISTRY)**

**GEGEWENS VIR FISIESTE WETENSKAPPE GRAAD 12
VRAESTEL 2 (CHEMIE)**

TABLE 1: PHYSICAL CONSTANTS/TABEL 1: FISIESTE KONSTANTES

NAME/NAAM	SYMBOL/SIMBOOL	VALUE/WAARDE
Standard pressure <i>Standaarddruk</i>	p^\ominus	$1,013 \times 10^5 \text{ Pa}$
Molar gas volume at STP <i>Molêre gasvolume by STD</i>	V_m	$22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$
Standard temperature <i>Standaardtemperatuur</i>	T^\ominus	273 K

TABLE 2: FORMULAE/TABEL 2: FORMULES

$n = \frac{m}{M}$	$c = \frac{n}{V}$ <p style="text-align: center;"><i>or/of</i></p> $c = \frac{m}{MV}$
$q = I \Delta t$ $W = Vq$	$E_{\text{cell}}^\ominus = E_{\text{cathode}}^\ominus - E_{\text{anode}}^\ominus / E_{\text{sel}}^\ominus = E_{\text{katode}}^\ominus - E_{\text{anode}}^\ominus$ <p style="text-align: center;"><i>or/of</i></p> $E_{\text{cell}}^\ominus = E_{\text{reduction}}^\ominus - E_{\text{oxidation}}^\ominus / E_{\text{sel}}^\ominus = E_{\text{reduksie}}^\ominus - E_{\text{oksidasie}}^\ominus$ <p style="text-align: center;"><i>or/of</i></p> $E_{\text{cell}}^\ominus = E_{\text{oxidising agent}}^\ominus - E_{\text{reducing agent}}^\ominus /$ $E_{\text{sel}}^\ominus = E_{\text{oksideermiddel}}^\ominus - E_{\text{reduseermiddel}}^\ominus$

TABLE 3: THE PERIODIC TABLE OF ELEMENTS
TABEL 3: DIE PERIODIEKE TABEL VAN ELEMENTE

1 (I)	2 (II)	3	4	5	6	7	8	9	10	11	12	13 (III)	14 (IV)	15 (V)	16 (VI)	17 (VII)	18 (VIII)
2,1 1 H																	2 He 4
1,0 3 Li	1,5 4 Be											2,0 5 B	2,5 6 C	3,0 7 N	3,5 8 O	4,0 9 F	10 Ne 20
0,9 11 Na	1,2 12 Mg											1,5 13 Al	1,8 14 Si	2,1 15 P	2,5 16 S	3,0 17 Cl	18 Ar 40
0,8 19 K	1,0 20 Ca	1,3 21 Sc	1,5 22 Ti	1,6 23 V	1,7 24 Cr	1,8 25 Mn	1,9 26 Fe	2,0 27 Co	2,1 28 Ni	1,9 29 Cu 63,5	1,6 30 Zn	1,6 31 Ga	1,8 32 Ge	2,0 33 As	2,4 34 Se	2,8 35 Br	36 Kr 84
0,8 37 Rb	1,0 38 Sr	1,2 39 Y	1,4 40 Zr	1,5 41 Nb	1,8 42 Mo	1,9 43 Tc	2,2 44 Ru	2,2 45 Rh	2,2 46 Pd	1,9 47 Ag	1,7 48 Cd	1,7 49 In	1,8 50 Sn	1,9 51 Sb	2,1 52 Te	2,5 53 I	54 Xe 131
0,7 55 Cs	0,9 56 Ba	57 La	1,6 72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	1,8 81 Tl	1,8 82 Pb	1,9 83 Bi	2,0 84 Po	2,5 85 At	86 Rn
0,7 87 Fr	0,9 88 Ra 226	89 Ac															
			58 Ce 140	59 Pr 141	60 Nd 144	61 Pm	62 Sm 150	63 Eu 152	64 Gd 157	65 Tb 159	66 Dy 163	67 Ho 165	68 Er 167	69 Tm 169	70 Yb 173	71 Lu 175	
			90 Th 232	91 Pa	92 U 238	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr	

KEY/SLEUTEL

Atomic number
Atoomgetal
Symbol
Simbool

Electronegativity
Elektronegatiwiteit

1,9	29
	Cu
	63,5

Approximate relative atomic mass
Benaderde relatiewe atoommassa

TABLE 4A: STANDARD REDUCTION POTENTIALS
TABEL 4A: STANDAARD- REDUKSIEPOTENSIALE

Increasing oxidising ability/Toenemende oksiderende vermoë

Increasing reducing ability/Toenemende reduserende vermoë

Half-reactions/Halfreaksies	E^{\ominus} (V)
$F_2(g) + 2e^- \rightleftharpoons 2F^-$	+ 2,87
$Co^{3+} + e^- \rightleftharpoons Co^{2+}$	+ 1,82
$H_2O_2 + 2H^+ + 2e^- \rightleftharpoons 2H_2O$	+ 1,77
$MnO_4^- + 8H^+ + 5e^- \rightleftharpoons Mn^{2+} + 4H_2O$	+ 1,51
$Ba^{2+} + 2e^- \rightleftharpoons Ba$	- 2,90
$Cl_2(g) + 2e^- \rightleftharpoons 2Cl^-$	+ 1,36
$Cr_2O_7^{2-} + 14H^+ + 6e^- \rightleftharpoons 2Cr^{3+} + 7H_2O$	+ 1,33
$Na^+ + e^- \rightleftharpoons Na$	- 2,71
$O_2(g) + 4H^+ + 4e^- \rightleftharpoons 2H_2O$	+ 1,23
$MnO_2 + 4H^+ + 2e^- \rightleftharpoons Mn^{2+} + 2H_2O$	+ 1,23
$Mn^{3+} + e^- \rightleftharpoons Mn^{2+}$	+ 1,50
$Cr^{3+} + e^- \rightleftharpoons Cr^{2+}$	+ 0,01
$Br_2(l) + 2e^- \rightleftharpoons 2Br^-$	+ 1,07
$NO_3^- + 4H^+ + 3e^- \rightleftharpoons NO(g) + 2H_2O$	+ 0,96
$Zn^{2+} + 2e^- \rightleftharpoons Zn$	- 0,76
$Hg^{2+} + 2e^- \rightleftharpoons Hg(l)$	+ 0,85
$Fe^{3+} + e^- \rightleftharpoons Fe^{2+}$	+ 0,77
$NO_3^- + 2H^+ + e^- \rightleftharpoons NO_2(g) + H_2O$	+ 0,80
$Cd^{2+} + 2e^- \rightleftharpoons Cd$	- 0,40
$O_2(g) + 2H^+ + 2e^- \rightleftharpoons H_2O$	+ 1,23
$Co^{3+} + e^- \rightleftharpoons Co^{2+}$	+ 1,82
$Ni^{2+} + 2e^- \rightleftharpoons Ni$	- 0,25
$Sn^{4+} + 2e^- \rightleftharpoons Sn^{2+}$	+ 0,15
$Pb^{2+} + 2e^- \rightleftharpoons Pb$	- 0,13
$SO_2 + 4H^+ + 4e^- \rightleftharpoons S + 2H_2O$	+ 0,17
$2H_2O + 2e^- \rightleftharpoons H_2(g) + 2OH^-$	- 0,83
$S + 2H^+ + 2e^- \rightleftharpoons H_2S(g)$	+ 0,14
$SO_4^{2-} + 4H^+ + 2e^- \rightleftharpoons SO_2(g) + 2H_2O$	+ 0,17
$Cu^{2+} + 2e^- \rightleftharpoons Cu$	+ 0,16
$SO_4^{2-} + 4H^+ + 2e^- \rightleftharpoons SO_2(g) + 2H_2O$	+ 0,17
$S + 2H^+ + 2e^- \rightleftharpoons H_2S(g)$	+ 0,14
$2H_2O + 2e^- \rightleftharpoons H_2(g) + 2OH^-$	- 0,83
$SO_2 + 4H^+ + 4e^- \rightleftharpoons S + 2H_2O$	+ 0,17
$Pb^{2+} + 2e^- \rightleftharpoons Pb$	- 0,13
$Cu^{2+} + 2e^- \rightleftharpoons Cu$	+ 0,16
$Sn^{4+} + 2e^- \rightleftharpoons Sn^{2+}$	+ 0,15
$O_2(g) + 2H^+ + 2e^- \rightleftharpoons H_2O$	+ 1,23
$Ce^{4+} + e^- \rightleftharpoons Ce^{3+}$	+ 1,61
$Cd^{2+} + 2e^- \rightleftharpoons Cd$	- 0,40
$NO_3^- + 2H^+ + e^- \rightleftharpoons NO_2(g) + H_2O$	+ 0,80
$Fe^{3+} + e^- \rightleftharpoons Fe^{2+}$	+ 0,77
$Hg^{2+} + 2e^- \rightleftharpoons Hg(l)$	+ 0,85
$NO_3^- + 4H^+ + 3e^- \rightleftharpoons NO(g) + 2H_2O$	+ 0,96
$2H_2O + 2e^- \rightleftharpoons H_2(g) + 2OH^-$	- 0,83
$Br_2(l) + 2e^- \rightleftharpoons 2Br^-$	+ 1,07
$Cl_2 + 2e^- \rightleftharpoons 2Cl^-$	+ 1,36
$Pt^{2+} + 2e^- \rightleftharpoons Pt$	+ 1,20
$Mn^{3+} + e^- \rightleftharpoons Mn^{2+}$	+ 1,50
$MnO_2 + 4H^+ + 2e^- \rightleftharpoons Mn^{2+} + 2H_2O$	+ 1,23
$O_2(g) + 4H^+ + 4e^- \rightleftharpoons 2H_2O$	+ 1,23
$Mg^{2+} + 2e^- \rightleftharpoons Mg$	- 2,38
$Cr_2O_7^{2-} + 14H^+ + 6e^- \rightleftharpoons 2Cr^{3+} + 7H_2O$	+ 1,33
$Ca^{2+} + 2e^- \rightleftharpoons Ca$	- 2,87
$Cl_2(g) + 2e^- \rightleftharpoons 2Cl^-$	+ 1,36
$MnO_4^- + 8H^+ + 5e^- \rightleftharpoons Mn^{2+} + 4H_2O$	+ 1,51
$Ba^{2+} + 2e^- \rightleftharpoons Ba$	- 2,90
$H_2O_2 + 2H^+ + 2e^- \rightleftharpoons 2H_2O$	+ 1,77
$Co^{3+} + e^- \rightleftharpoons Co^{2+}$	+ 1,82
$F_2(g) + 2e^- \rightleftharpoons 2F^-$	+ 2,87

TABLE 4B:

STANDARD REDUCTION POTENTIALS
TABEL 4B: STANDAARD- REDUKSIEPOTENSIALE

Increasing oxidising ability/Toenemende oksiderende vermoë



Increasing reducing ability/Toenemende reduserende vermoë

